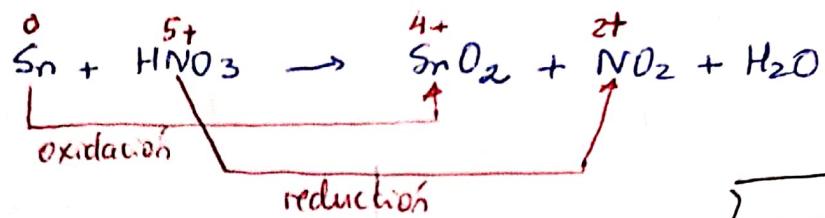


10 g carbio 120 M  
92,5% HgS

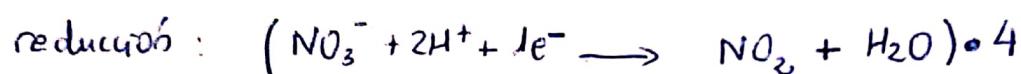
$$10 \text{ g carbio} \cdot \frac{92,5 \text{ g HgS}}{100 \text{ g carbio}} \cdot \frac{1 \text{ mol HgS}}{232,67 \text{ g HgS}} \cdot \frac{8 \text{ mol HNO}_3}{3 \text{ mol HgS}} \cdot \frac{1 \text{ L dis HNO}_3}{12,0 \text{ mol HNO}_3} \cdot \frac{10^3 \text{ mL}}{1 \text{ L}}$$

$$= \boxed{8,8 \text{ mL dis HNO}_3}$$

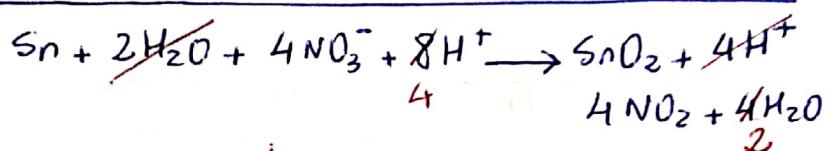
18 FLECHAS



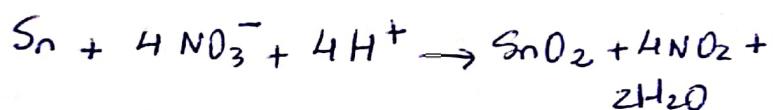
α)



FORMA  
IONICA



$\downarrow$  Simplifico  $H^+$  y  $H_2O$



NOTA: Cojo la forma molecular de la especie ( $\text{SnO}_2$ ) y no la iónica ( $\text{Sn}^{4+}$ )

## RECORDATORIO:

Forma iônica → Ácidos  
 $(H_4)$   
 → BASES  
 $(NaOH)$   
 ↓ SALES  
 $(FeCl_3)$

## FORMA MOLECULAR



b)

$2g$       V?  
16%  
 $1,09 g/mL$

