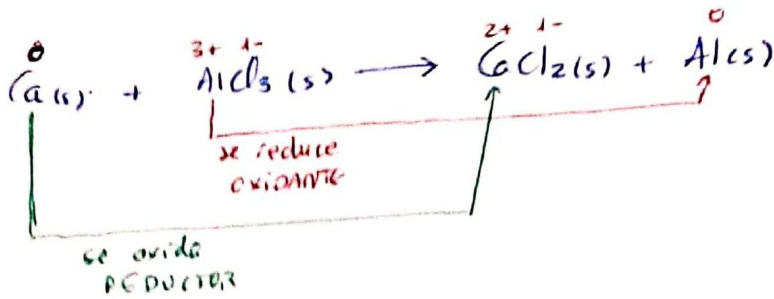
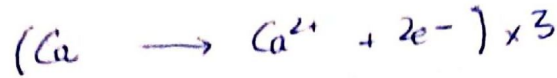


7a)



OXIDACIÓN



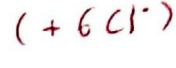
REDUCCIÓN



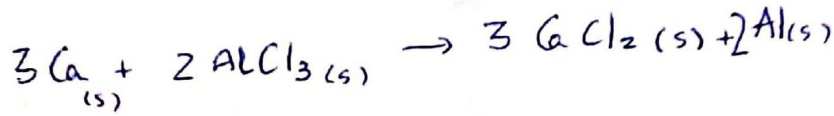
Ecuación en forma iónica ajustada



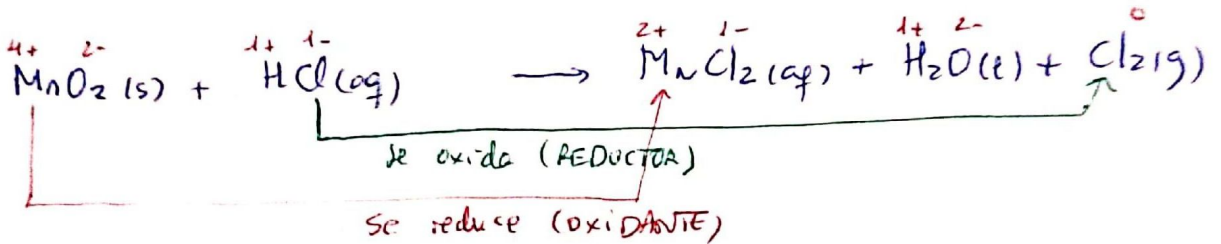
en ambos miembros por tanto



FORMA MOLECULAR



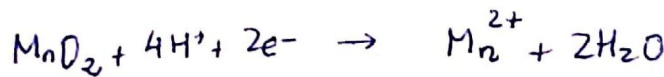
7b)



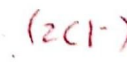
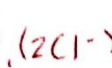
OXIDACIÓN



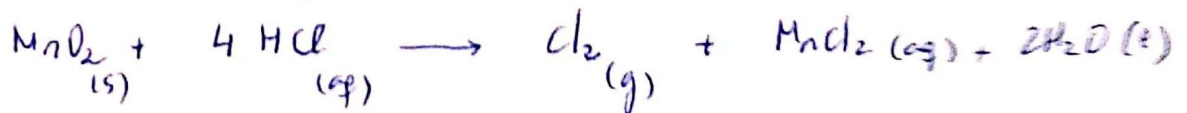
REDUCCIÓN



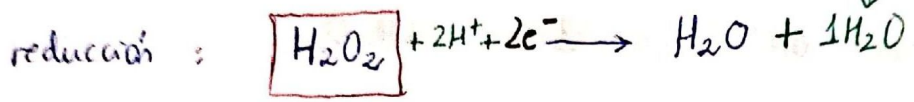
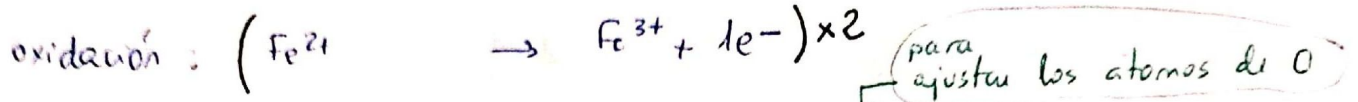
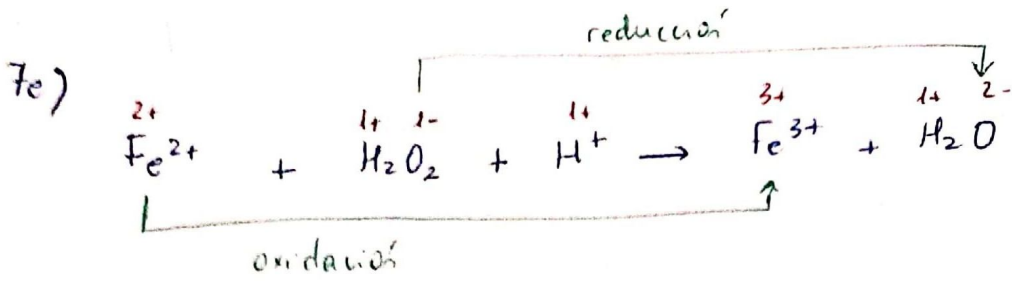
Ecuación prima iónica



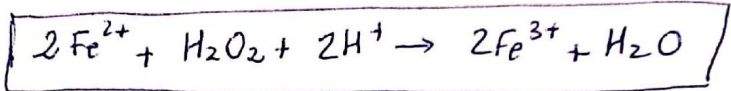
FORMA MOLECULAR





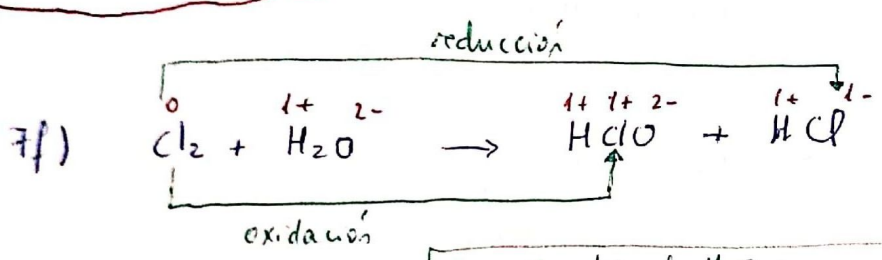


• PODEMOS ESCRIBIR LA FORMA MOLECULAR ( $\text{H}_2\text{O}_2$ )  
 O LA FORMA IÓNICA ( $\text{O}_2^{2-}$ )



↳ No hay que "traducir" a forma molecular.

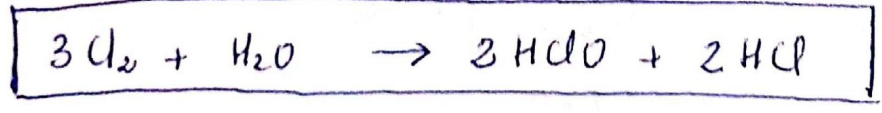
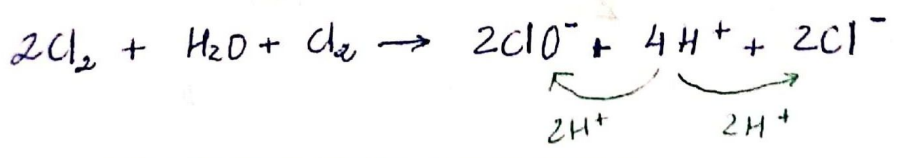
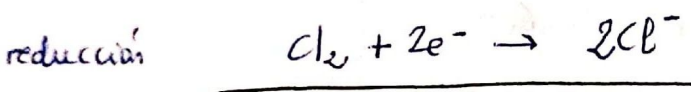
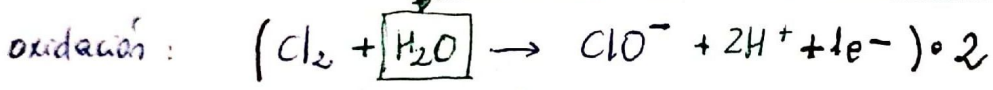
Ya la estoy expresando de modo similar al enunciado



DISMUTACIÓN O DESPROPORCIÓN  
 El  $\text{Cl}_2$  se oxida y reduce a la vez  
 $\text{HClO} \rightarrow \text{H}^+ + \text{ClO}^-$

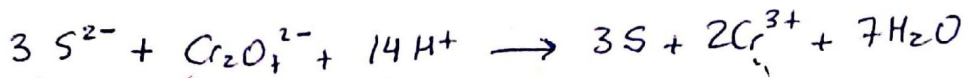
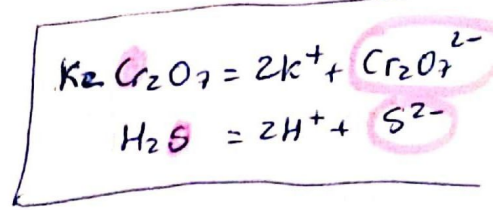
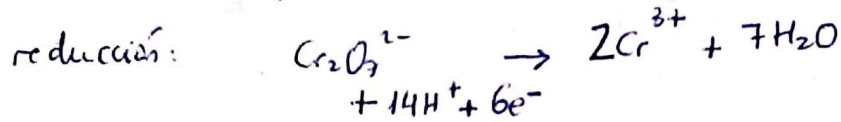
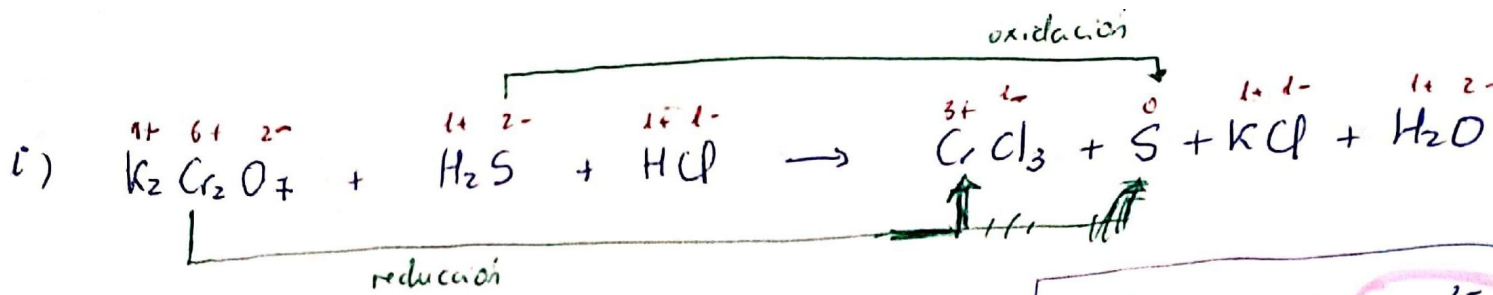
↳ FORMA IÓNICA

sumo 1 molec de  $\text{H}_2\text{O}$  para ajustar el O

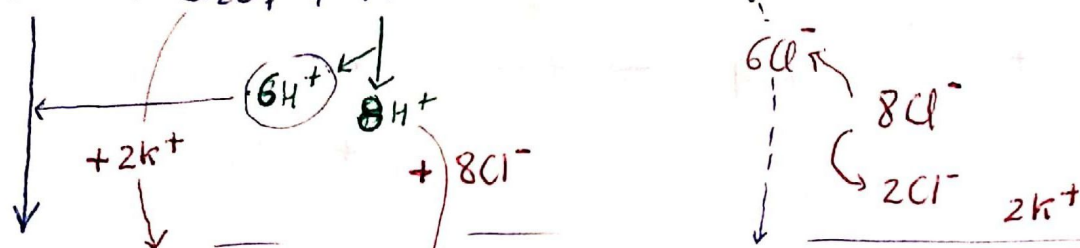


Forma iónica  
 ↓ "traducir"  
 Forma molecular

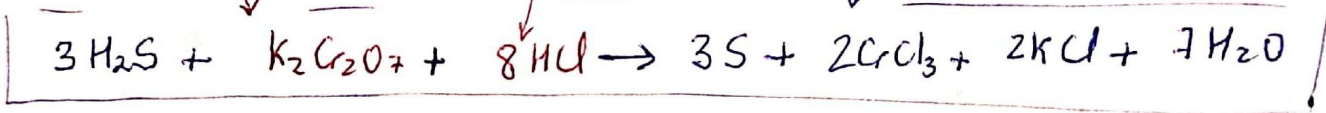




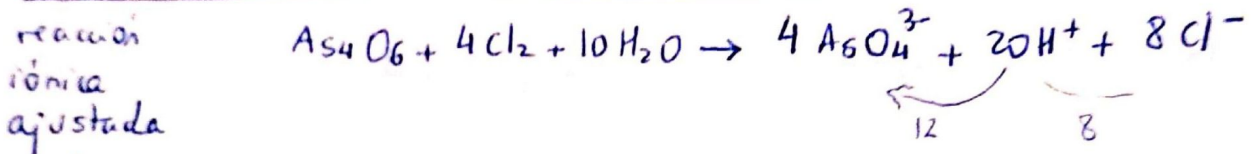
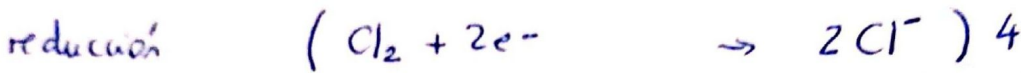
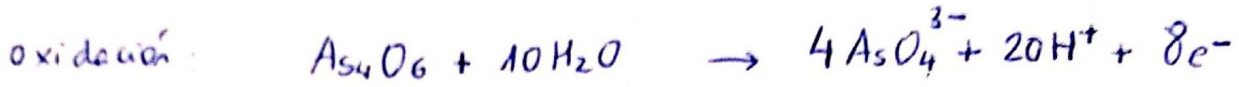
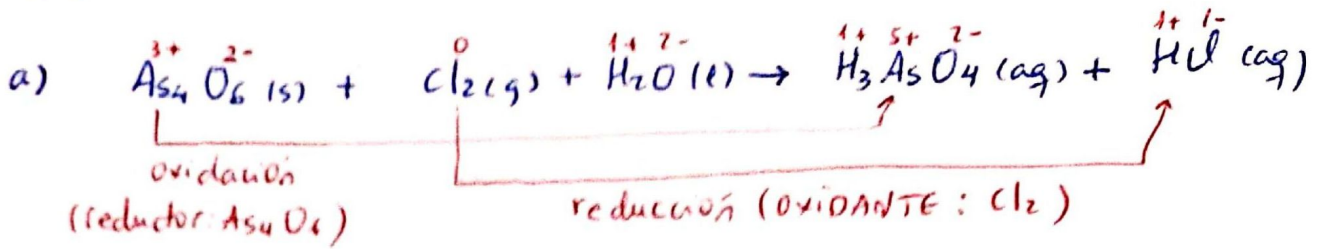
Forma iónica



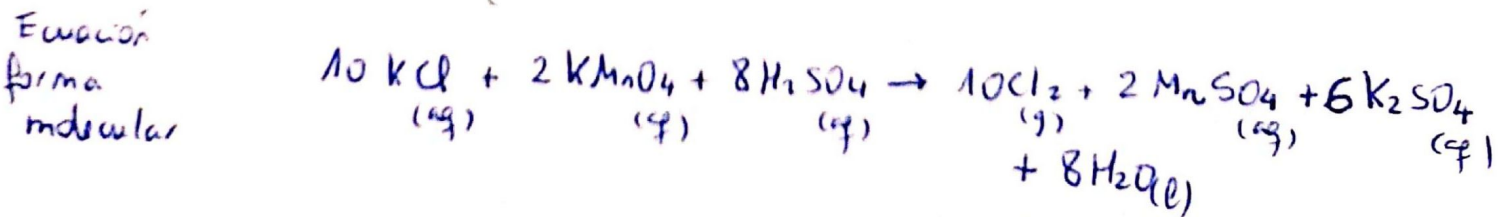
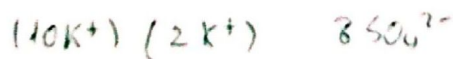
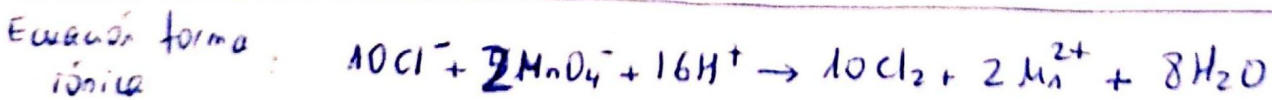
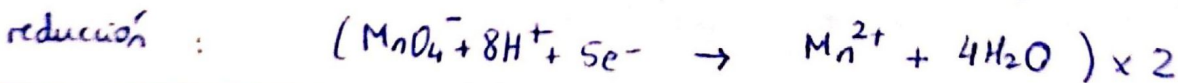
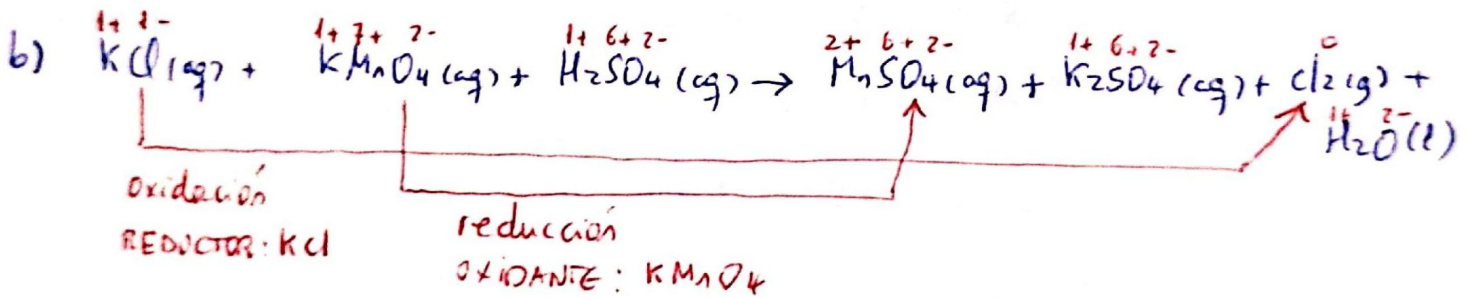
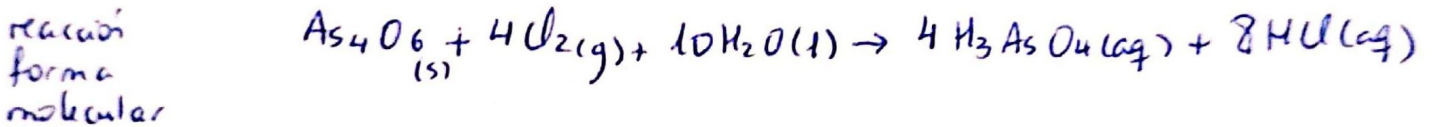
Forma molecular



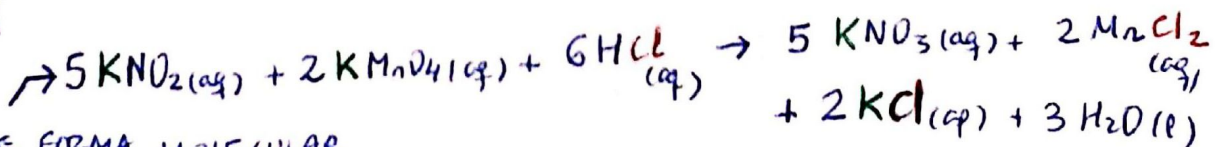
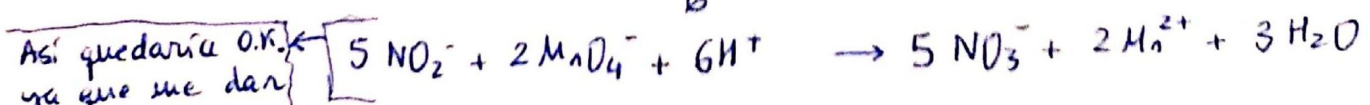
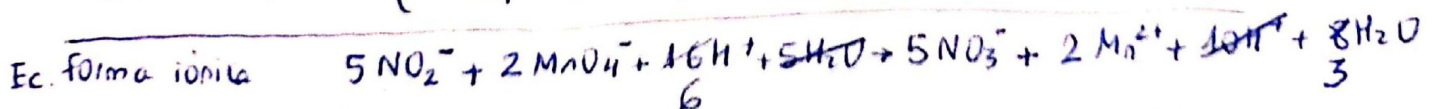
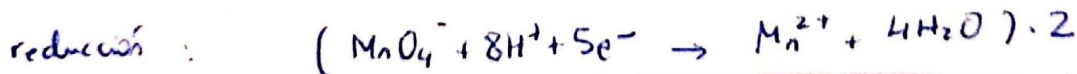
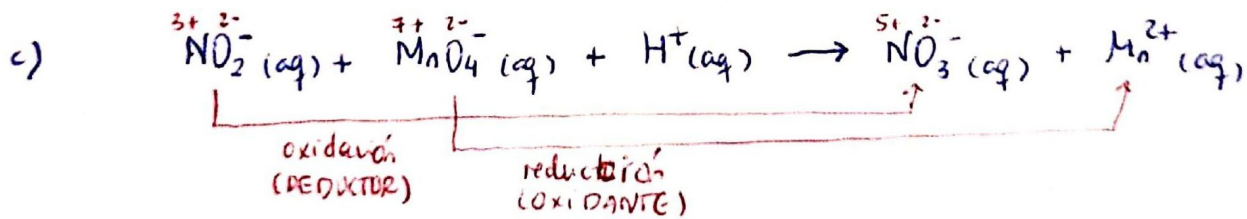
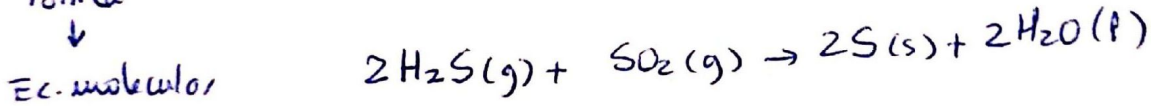
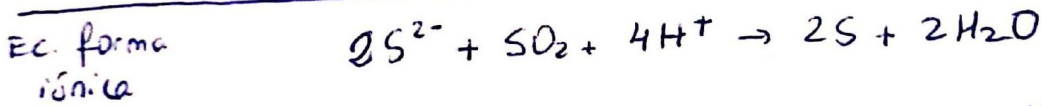
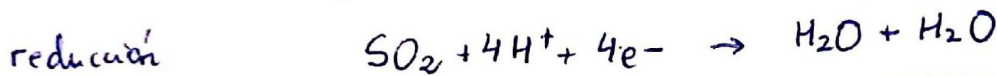
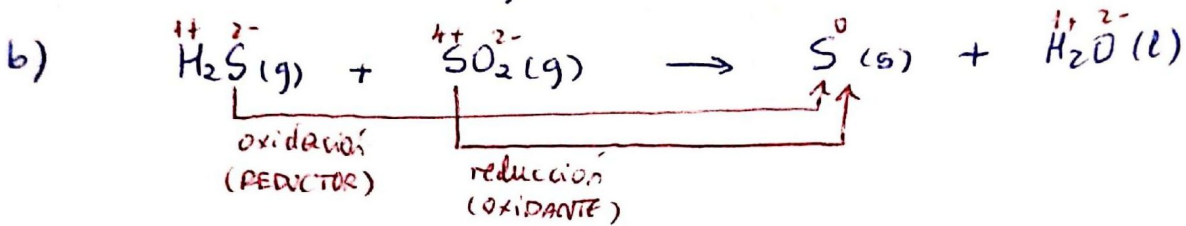
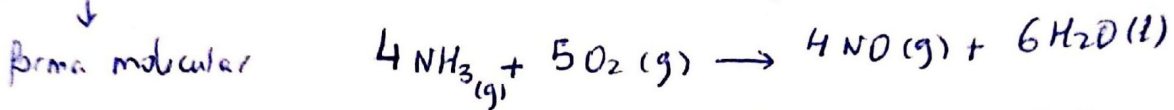
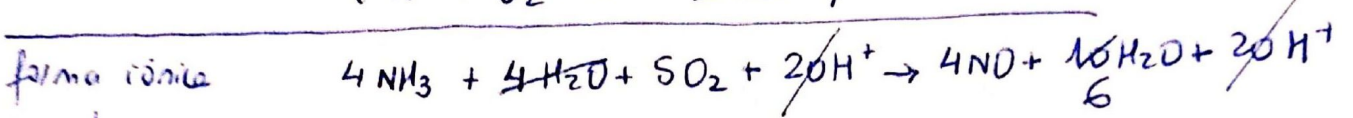
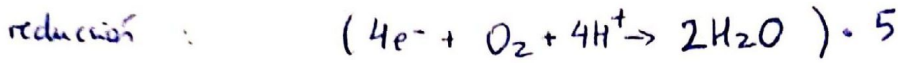
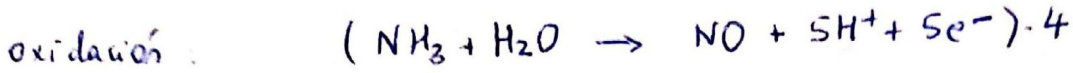
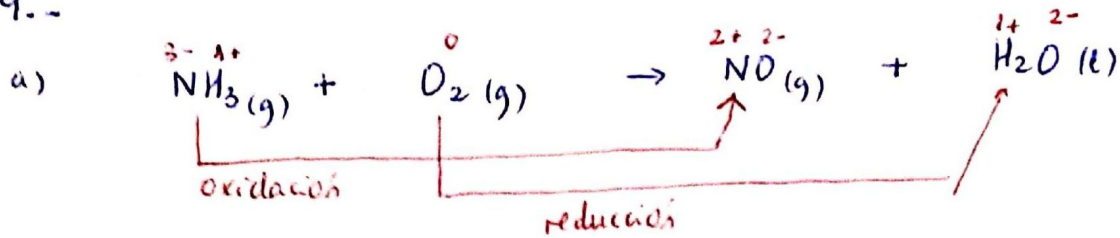
8. -



↓



9.-



UNA POSIBLE FORMA MOLECULAR