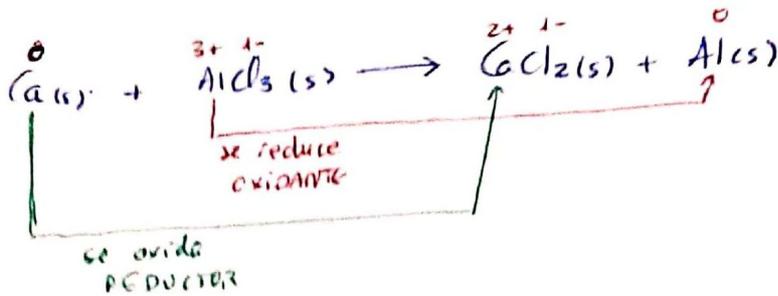
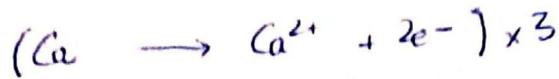


7a)



OXIDACIÓN



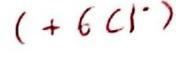
REDUCCIÓN



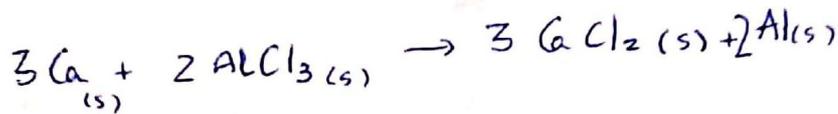
Ecuación en
forma iónica
ajustada



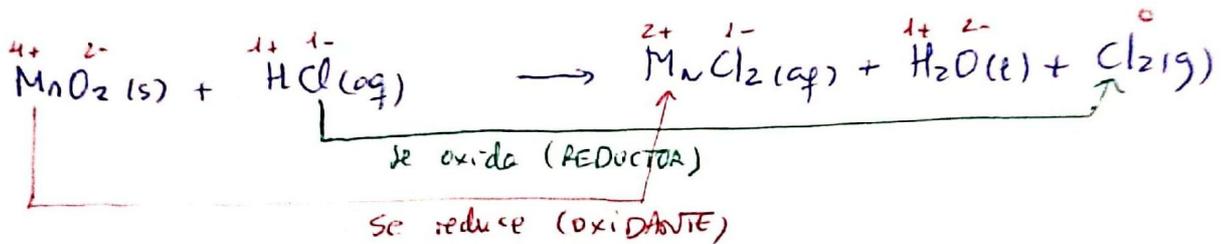
en ambos miembros
por tanto



FORMA
MOLECULAR



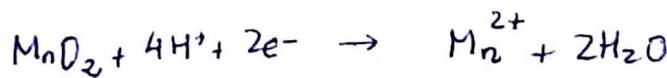
7b)



OXIDACIÓN



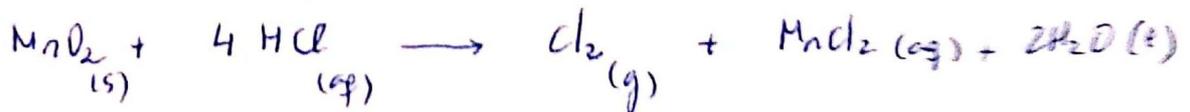
REDUCCIÓN

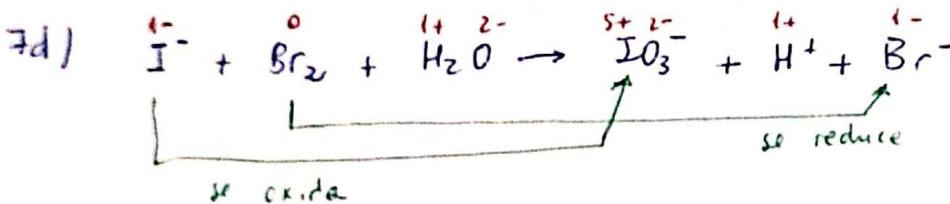
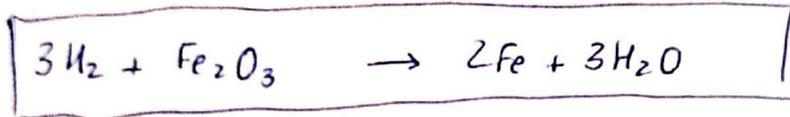
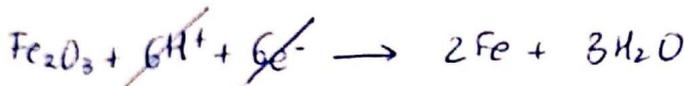
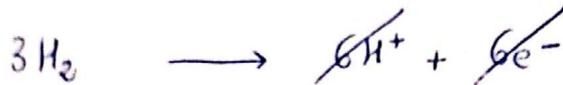
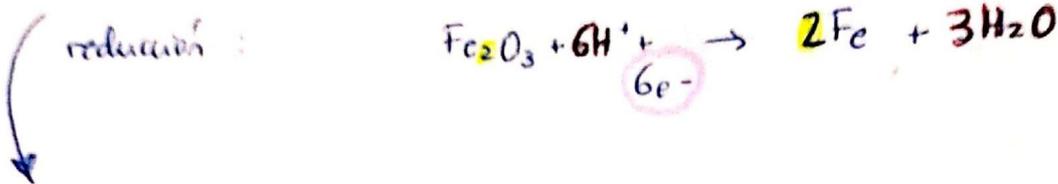
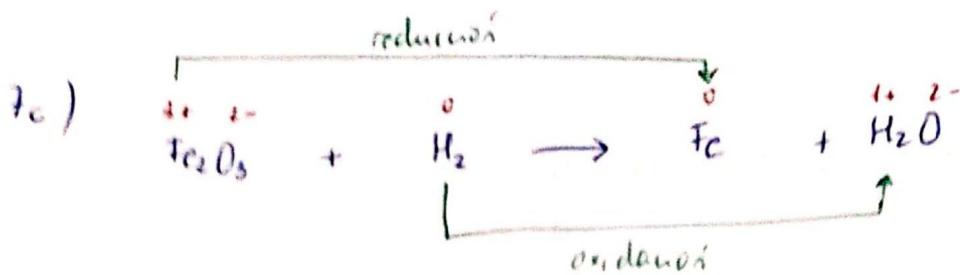


Ecuación
forma iónica

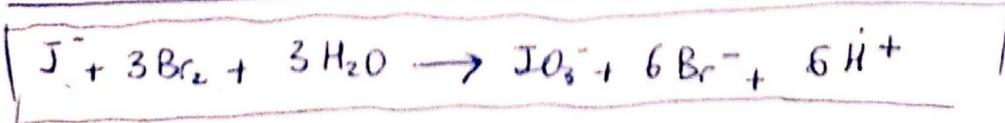
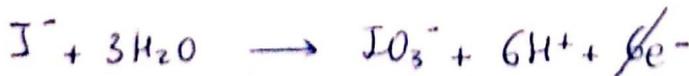


FORMA
MOLECULAR

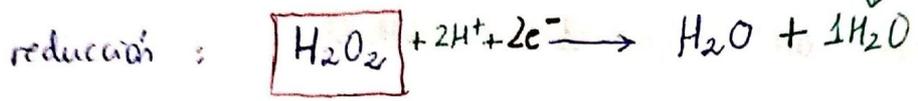
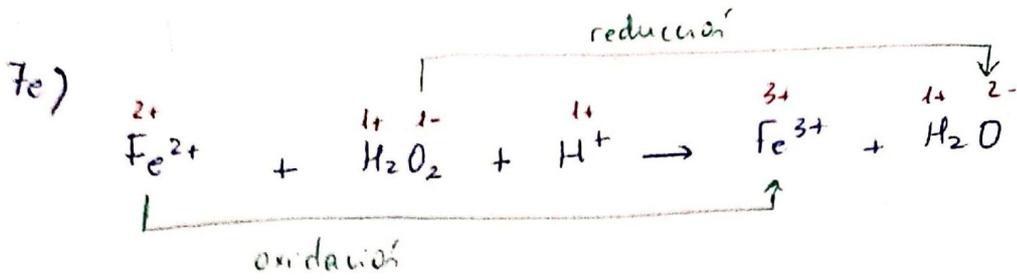




luego abajo

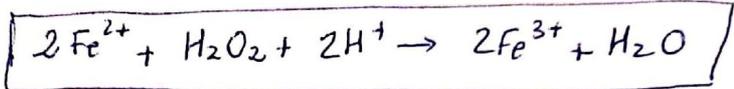


ME LA DABAN EN FORMA IÓNICA POR LO QUE NO HAY QUE PASARLA A FORMA MOLECULAR



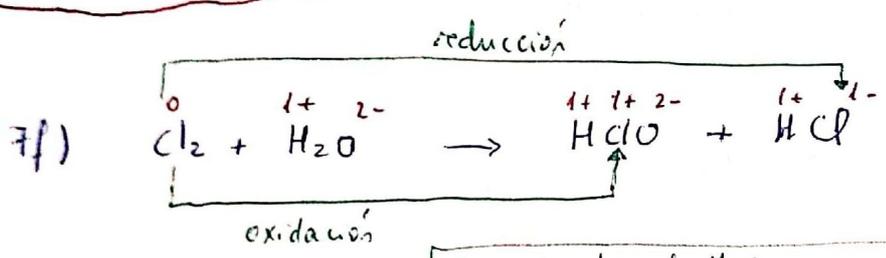
para ajustar los átomos de O

• PODEMOS ESCRIBIR LA FORMA MOLECULAR (H₂O₂)
O LA FORMA IÓNICA (O₂²⁻)



↳ No hay que "traducir" a forma molecular.

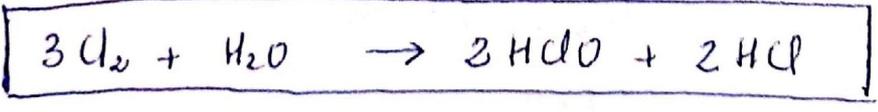
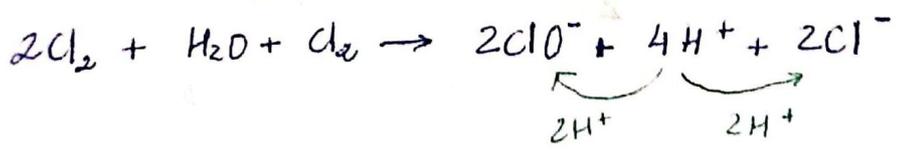
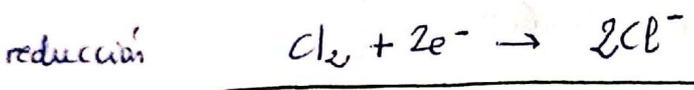
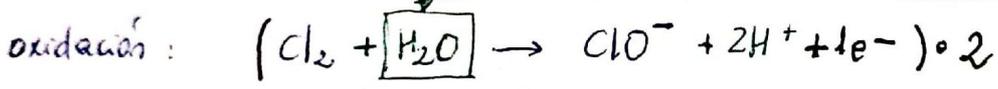
Ya la estoy expresando de modo similar al enunciado



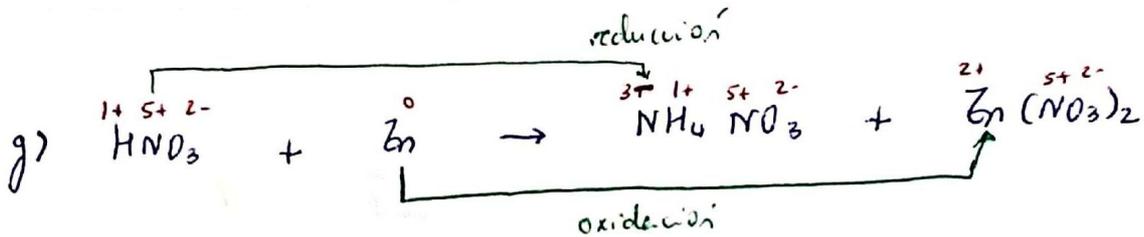
DISMUTACIÓN O DESPROPORCIÓN
El Cl₂ se oxida y reduce a la vez
 $\text{HClO} \rightarrow \text{H}^+ + \text{ClO}^-$

↳ FORMA IÓNICA

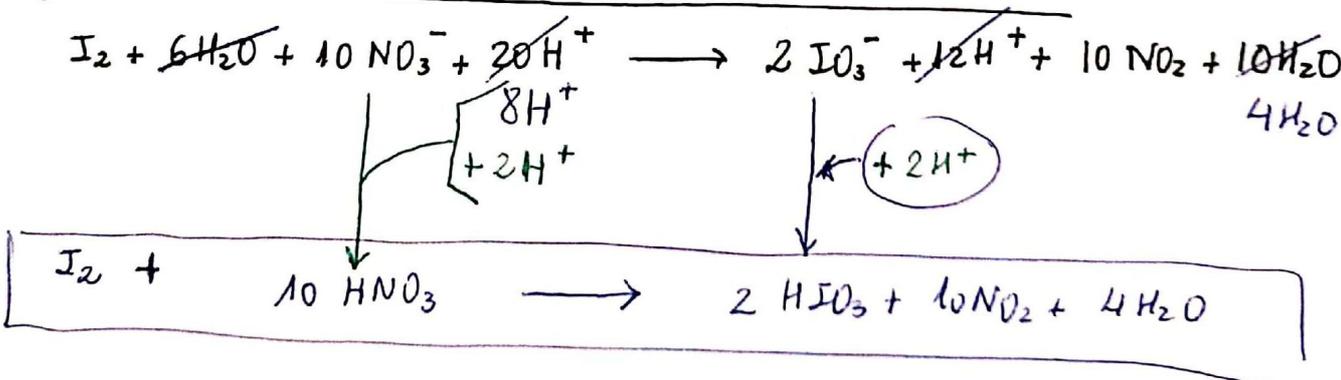
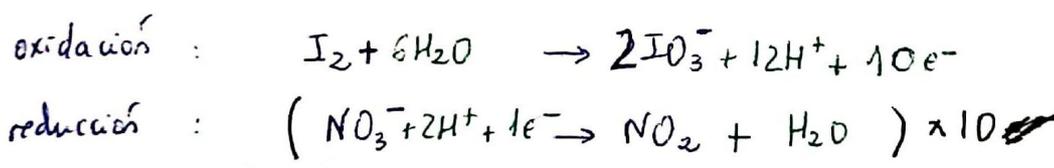
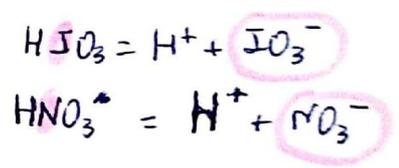
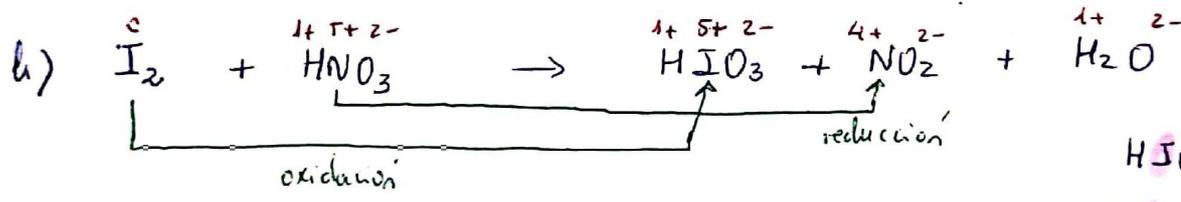
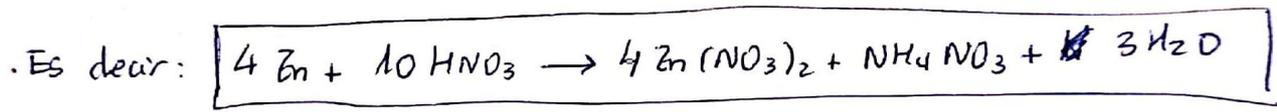
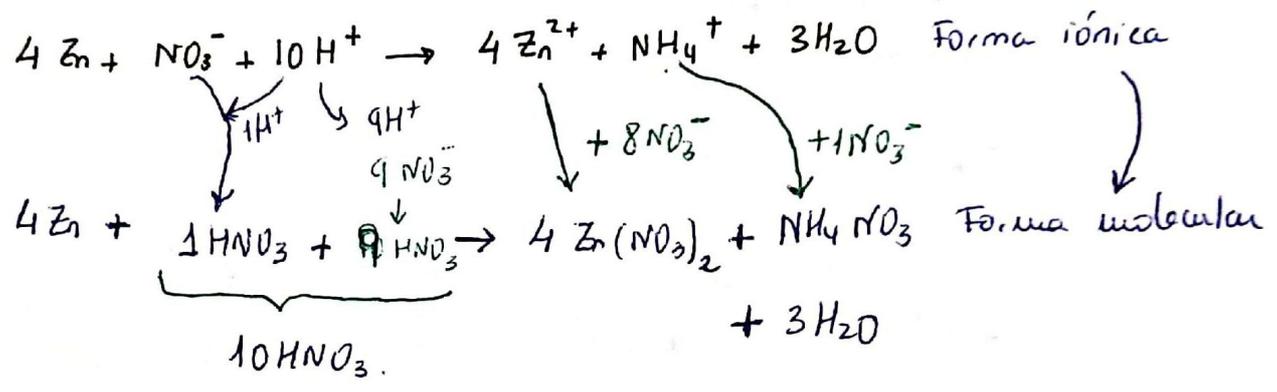
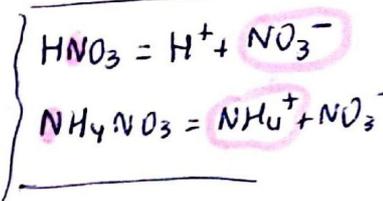
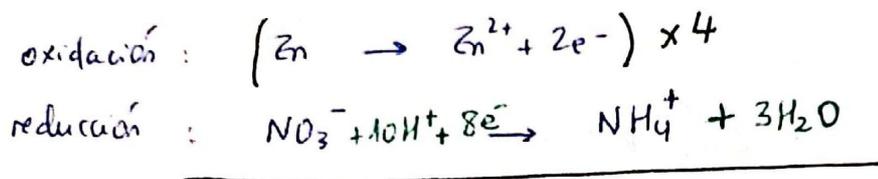
sumo 1 molec de H₂O para ajustar el O

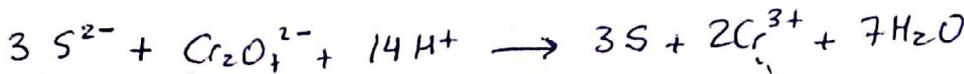
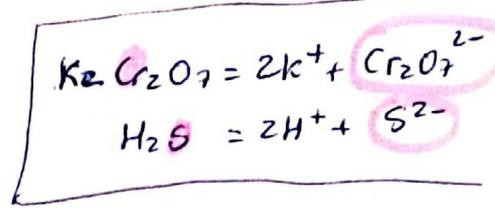
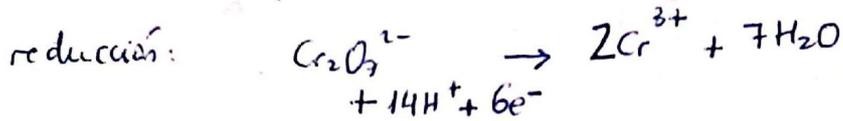
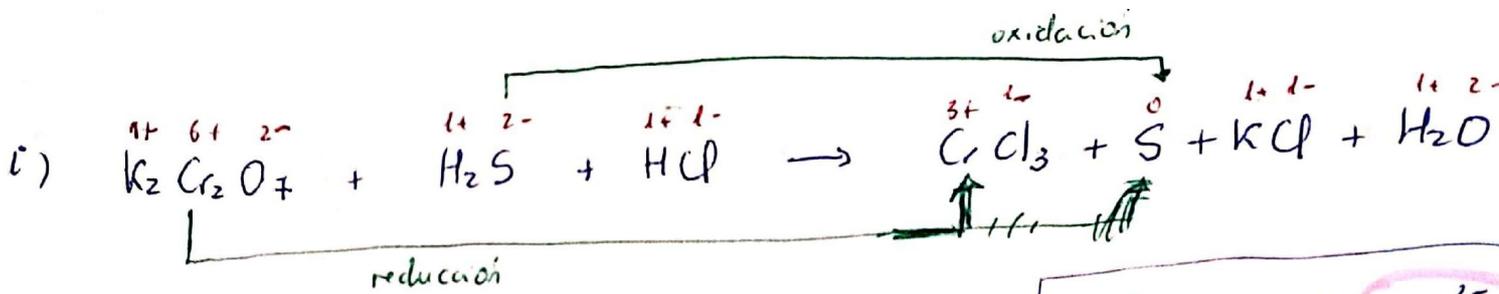


Forma iónica
↓ "traducir"
Forma molecular



$(\overset{-3 \quad -1 \quad +}{\text{NH}_4})^+$
 es más EN
 pero se escribe
 a la izda

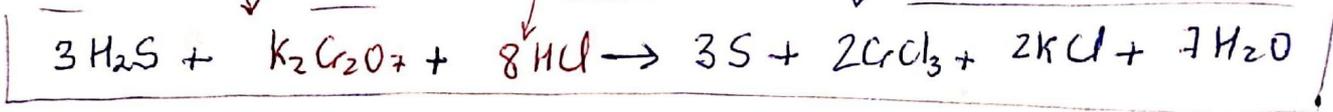




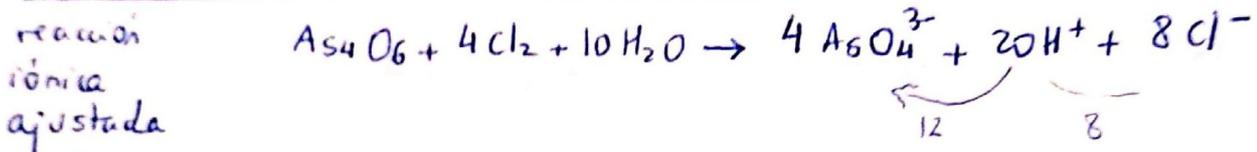
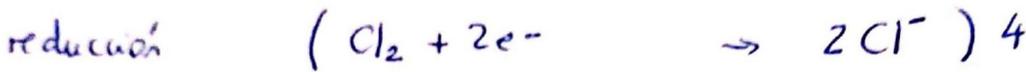
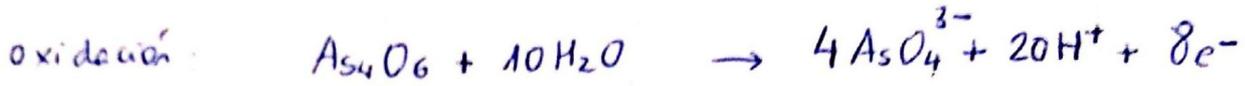
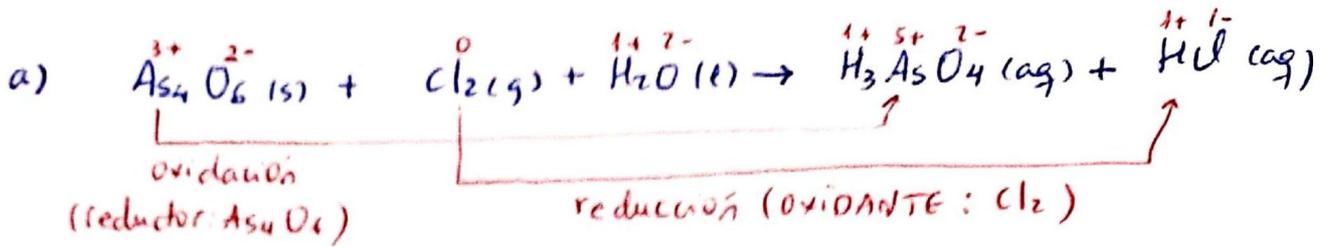
Forma iónica



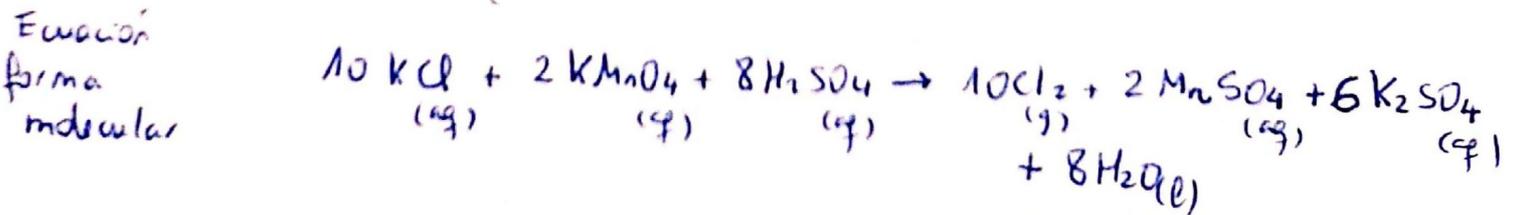
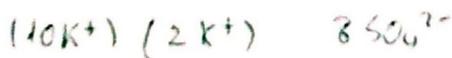
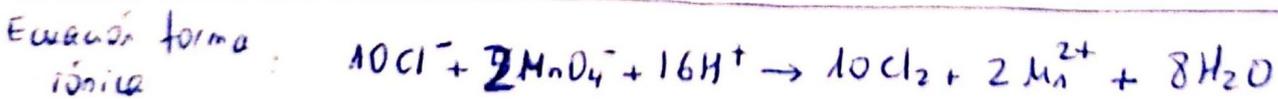
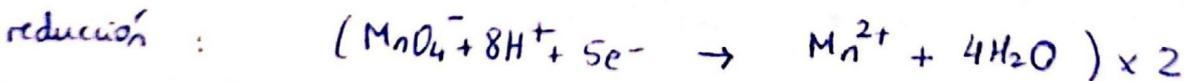
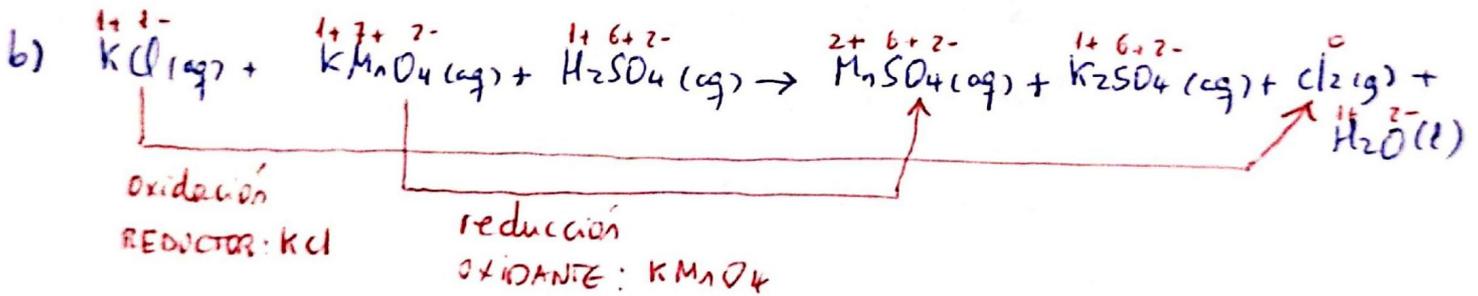
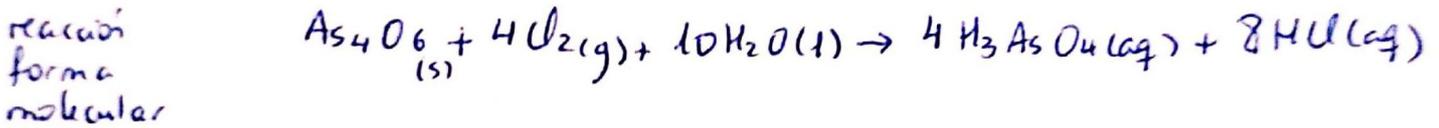
Forma molecular



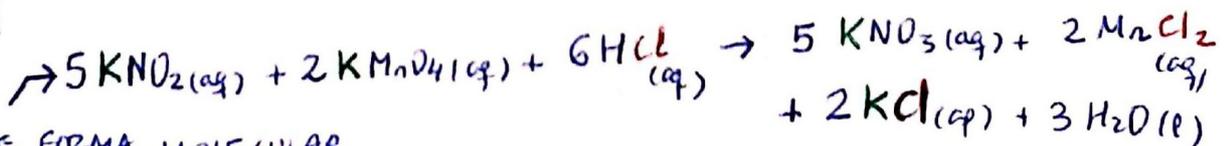
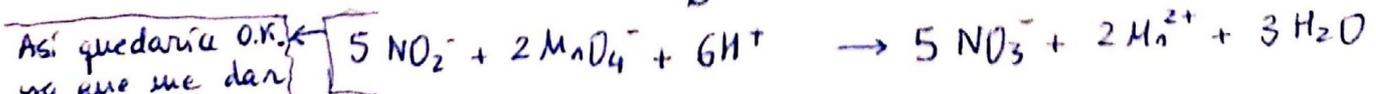
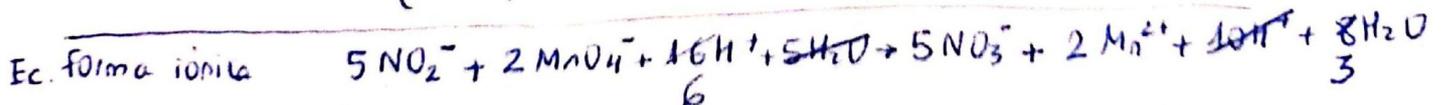
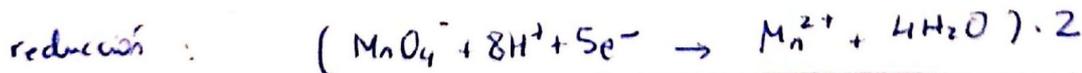
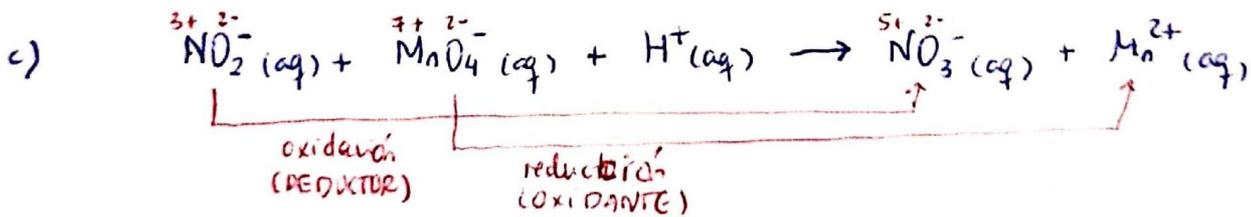
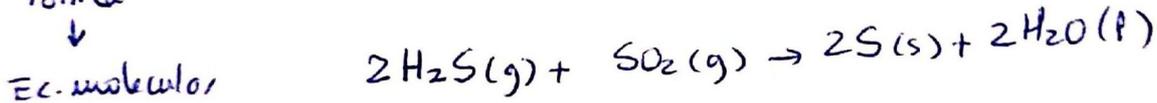
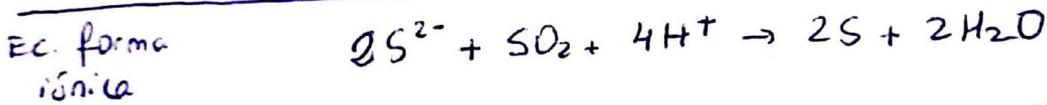
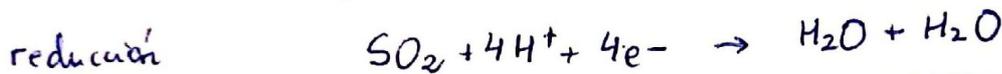
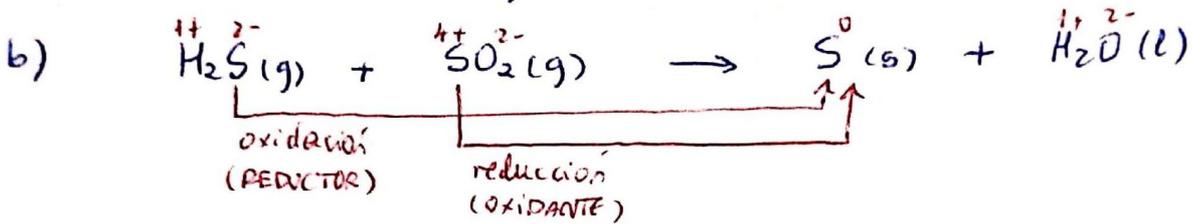
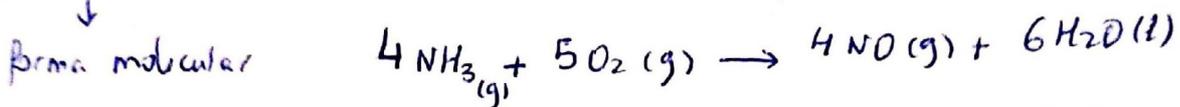
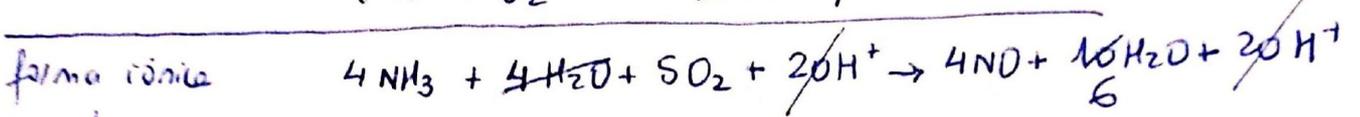
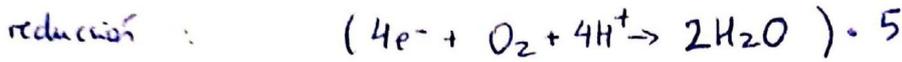
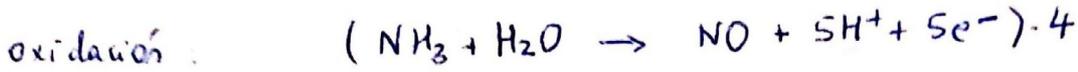
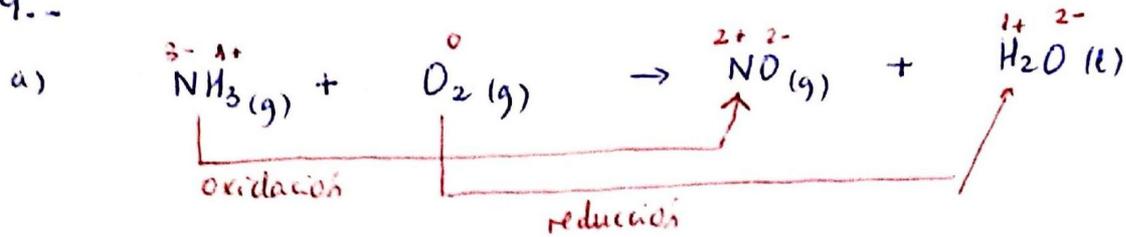
8. -



↓



9.-



UNA POSIBLE FORMA MOLECULAR